

This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1921 Excerpt: ...per cent, of nitrogen respectively, the remainder being copper, with the exception of about 0.7 per cent, of hydrogen.-which was probably present in the occluded condition. Beilby and Henderson state that no nitride is formed unless the ammonia is present in large excess, the maximum percentage of nitrogen fixed in their experiments being 45. Guntz and Bassett prepared the nitride from cuprous instead of cupric oxide, t The reaction begins at 250°C. and is rapid at 3100, but the nitride begins to decompose at the temperature of production. The products obtained contained about 6 per cent, of nitrogen, and it was found possible to remove the free copper by treatment with a solution of ammonia and ammonium carbonate. The nitride thus purified contained about 68 per cent, of nitrogen. The formula Cu_3N , corresponding to lithium nitride, requires 686 per cent, Copper nitride may also be obtained by the interaction of copper oxide with magnesium nitride, but the product is very impure. Similarly, on heating cuprous or cupric chloride with lithium nitride a reaction takes place. This is, however, so violent that copper only is obtained, the nitride being decomposed by the high temperature. Cuprous nitride is a greenish-black powder, which decomposes into its elements on being heated to 3000 C Schrotter states that the decomposition is accompanied by a red glow. It dissolves in acids, forming ammonium salts and cuprous salts: $2Cu_3N + 8HCl = 3Cu_2Cl_2 + 2NH_4Cl$. Cuprous nitride is decomposed normally, although Trans. Chem. Soc, 1901, 71, 1252. t Bull. Soc. Chim. (3), 1906, 35, 201. t Smits, Rec. Trav. Chim. Pay-Bas., 1896, 15, 135. not very quickly, by boiling caustic soda solution, with evolution of ammonia: $2Cu_3N + 3H_2O = 3Cu_2O + 2NH_3$. Silver Nitride. (Ag₃N). This bod...

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